The gas-phase chemiionization reaction between samarium and oxygen atoms: A theoretical study

PAULOVIC, Jozef, et al.

Abstract

The Sm + O chemiionization reaction has been investigated theoretically using a method that allows for correlation and relativistic effects. Potential energy curves have been calculated for several electronic states of SmO and SmO+. Comparison with available spectroscopic and thermodynamic values for these species is reported and a mechanism for the chemiionization reaction Sm + O is proposed. The importance of spin–orbit coupling in the excited states of SmO, in allowing this chemiionization reaction to take place, has been revealed by these calculations. This paper shows the metal-plus-oxidant chemiionization reaction.

Reference


DOI: 10.1063/1.1730034

Available at:
http://archive-ouverte.unige.ch/unige:3321

Disclaimer: layout of this document may differ from the published version.
The gas-phase chemiionization reaction between samarium and oxygen atoms: A theoretical study

Jozef Paulović
Department of Applied Chemistry, School of Engineering, The University of Tokyo, Tokyo 113-8656, Japan

Laura Gagliardi
Dipartimento di Chimica Fisica “F. Accascina,” Viale delle Scienze-Parco d’Orleans II, I-90128 Palermo, Italy

John M. Dyke
Department of Chemistry, University of Southampton, Highfield, Southampton SO17 1BJ, United Kingdom

Kimihiko Hirao
Department of Applied Chemistry, School of Engineering, The University of Tokyo, Tokyo 113-8656, Japan

(Received 23 December 2003; accepted 9 March 2004)

The Sm + O chemiionization reaction has been investigated theoretically using a method that allows for correlation and relativistic effects. Potential energy curves have been calculated for several electronic states of SmO and SmO⁺. Comparison with available spectroscopic and thermodynamic values for these species is reported and a mechanism for the chemiionization reaction Sm + O is proposed. The importance of spin–orbit coupling in the excited states of SmO, in allowing this chemiionization reaction to take place, has been revealed by these calculations. This paper shows the metal-plus-oxidant chemiionization reaction. © 2004 American Institute of Physics.

[I DOI: 10.1063/1.1730034]

I. INTRODUCTION

Chemiionization is an intriguing form of ionization in the gas-phase, which can occur between ground state reagents. It has been observed to occur in metal-plus-oxidant reactions and hydrocarbon–oxygen reactions in flames. subdivision

The simplest case is the reaction between a metal, M, and an oxygen atom, O, to give MO⁺ plus an electron. The reaction is exothermic, and can therefore be observed experimentally, if the dissociation energy of the metal oxide produced is greater than the adiabatic ionization energy of the neutral metal oxide. Unfortunately, although qualitative potential energy curves can be drawn to explain the production of ions and electrons for an M + O chemiionization reaction, subdivision no electronic structure calculations have yet been made for the states involved for any M + O reactions studied experimentally. This is almost certainly because most of the metals (M) for which this form of ionization has been observed are lanthanides and actinides, in which the MO and MO⁺ species are difficult to treat theoretically, not least because the metals have electronic configurations with unfilled f shells and give rise to a number of low-lying states in MO and MO⁺. Correlation, relativistic and spin–orbit effects must therefore be included.

This situation is in marked contrast with the chemiionization reactions O + CH(X₂H), O + CH(a⁻Σ⁻), O + CH₂C(ß²A’), and O + CH₂C(α⁻A₂), which have been identified in hydrocarbon flames. subdivision For each reaction, potential energy surfaces for the neutral and ion states involved have been computed (e.g., HCO and HCO⁺ in the case of O + CH) and the maximum electron kinetic energy for the chemiionization reaction calculated. subdivision These potential curves have allowed the mechanisms for the observed chemiionization processes to be established and the chemielectron bands to be assigned.

One of the simplest M + O chemiionization reactions for which a chemielectron spectrum has been recorded and the chemions (MO⁺) detected is the samarium plus oxygen atom reaction. Metal oxidation chemiionization reactions play important roles in the upper atmosphere, where collisions involving ground and excited metal atoms with oxidants, such as atomic oxygen, are a major source of ions. subdivision M + O chemiionization reactions, such as the Sm + O associative ionization reaction, have also been considered as a way of monitoring atomic oxygen in the upper atmosphere.

Although lanthanide monoxides (LnO) have been the subject of extensive theoretical and experimental studies, there are no ab initio studies on the electronic structures of SmO and SmO⁺. This is because of the complexity of the electronic structures of these species, in particular the high angular momentum and spin multiplicity of the low-lying states involved. These in turn derive from the low-lying electronic states of the samarium atom, which have partially filled 4f and 5d shells. Therefore a large number of molecular electronic states that are close in energy result, and these are further split and mixed by spin–orbit interaction. Hence, accurate calculations on SmO and SmO⁺ require incorporation of both relativistic and correlation effects to obtain reliable results.

This work calculates potential energy curves of SmO...
and SmO\(^+\), with the treatment of correlation, relativistic effects and spin–orbit coupling, to allow a comparison with available spectroscopic and thermodynamic values for these species and to allow a mechanism for the observed Sm+ O chemionization to be proposed. This work represents the first theoretical attempt to study the energetics and mechanism of a metal-plus-oxidant chemionization reaction.

II. COMPUTATIONAL DETAILS

The study was performed using the complete active space (CAS) SCF method\(^{16}\) with dynamic electron correlation added using multiconfigurational second order perturbation theory (CASPT2).\(^{15–17}\) Scalar relativistic effects were included via Douglas–Kroll (DK) Hamiltonian.\(^{18,19}\) The effects of spin–orbit (SO) coupling were calculated by using a newly developed method based on the CASSCF state interaction method (CASSI).\(^{20,21}\) Here, the CASSCF wave functions generated for a range of electronic states are allowed to mix under the influence of a spin–orbit Hamiltonian. The method has recently been described;\(^{22}\) see that article for details. It has been shown to be successful in the study of actinide compounds.\(^{23,24}\) All calculations were performed with a prerelease of the software MOLCAS 6.0.\(^{25}\)

A newly developed basis set of the ANO type was used for the Sm atom.\(^{26}\) The exponents were optimized, including scalar relativistic effects, through the use of the Douglas–Kroll Hamiltonian.\(^{18,19}\) A primitive set of 25s22p15d11f4g functions was contracted to 9s8p6d4f2g. For the O atom the ANO-L basis set\(^{28}\) of the MOLCAS library was used. A primitive set of 14s9p4d3f basis functions was contracted to 7s7p4d3f.

In the CASSCF treatment, the [Kr] 4d, 5s, 5p orbitals of Sm and 1s, 2s of O were kept doubly occupied. In total, 12 electrons were distributed over 13 active orbitals, i.e., 4f of Sm and 2p, 3p of O. In the subsequent MS-CASPT2 method, [Kr] 4d, 5s, 5p orbitals of Sm and 1s, 2s of O were frozen.

The potential energy curves for several electronic states of SmO and SmO\(^+\) were calculated at the CASPT2 level. Spectroscopic constants for the ground state of SmO and SmO\(^+\) were determined by using the program VIBROT, available in the MOLCAS 6.0 package. VIBROT fits the potential to an analytical form using cubic splines and computes the vibration and rotation constants for a diatomic molecule.

III. RESULTS

The ground states of SmO and SmO\(^+\) are 7\(\Delta\) and 9\(\Gamma\), respectively. The computed equilibrium spectroscopic constants (\(R_e\), \(D_e\), \(\omega_e\)) for the ground states for SmO and SmO\(^+\) are reported in Table I, together with the calculated first ionization energies (IEs) for SmO and Sm. The calculated values of \(D_e\) for SmO and SmO\(^+\) are within the quoted errors of the experimental values. Similar satisfactory agreement is obtained for the first IEs of Sm and SmO (see Table I). As mentioned in the Introduction, it is important to obtain reliable values for these \(D_e\)'s and IEs as a prerequisite to obtaining a reliable quantitative calculation of chemielectron energies. The computed \(R_e\) value for SmO (7\(\Delta\)) is too low (by ca. 0.03 Å) and the computed value of \(\omega_e\) is too high (by ca. 50 cm\(^{-1}\)). There are no experimental values available for SmO\(^+\) (6\(\Gamma\)).

The ground state electronic configurations of Sm and O atoms are \([\text{Xe}] 4f^66s^2\), \(7\text{F}_0\) and \([\text{He}] 2s^22p^4\), \(3\text{P}_2\), respectively. The electronic states of SmO are of the \(\Sigma, \Delta, \Pi, \Phi\) types. In \(C_2\) symmetry, the \(\Sigma, \Delta, \Phi\) states transform as irreducible representation 1 (irrep 1), while the \(\Pi, \Phi\) states transform as irreducible representation 2 (irrep 2). From the above atomic electronic configurations of Sm and O, the following electronic states of SmO will arise in \(C_2\): 11 septets, 11 quintets, and 11 nonets in irrep 1, and 10 septets, 10 quintets, and 10 nonets in irrep 2 (441 components in all). The potential energy curves for all these states have been calculated at the CASPT2 level of theory. In Fig. 1, the most significant spin-free curves of SmO are shown (i.e., the lower potential energy curves that are most relevant to the Sm+ O chemionization process), together with the spin-free potential energy curve for the ground state of SmO\(^+\).

At the equilibrium bond length, the ground state of SmO is 7\(\Delta\). The spin-free energy levels of SmO arising from the ground states of the Sm and O atoms lying within an energy range of 4.91 eV above the ground state at the equilibrium Sm–O bond length were allowed to mix under the spin–orbit

![Fig. 1. The spin-free potential energy curves for several states of SmO arising from the ground state of SmO\(^+\) (dotted line). The dotted horizontal line is the energy of Sm and O atoms at dissociation.](image-url)
Hamiltonian, and they give rise to more than 320 spin–orbit states. The first 12 spin–orbit states lie within an energy range of 1500 cm\(^{-1}\). A definite assignment of the \(\Omega\) value for such states is not possible, because of the small energy differences between the various states, which are far below the accuracy of the present quantum chemical method in determining electronic excitation energies, and the large mixing of the spin-free states. We will thus only compare the computed relative energies of the lowest spin states at the equilibrium bond length with the experimental values.\(^{29-31}\)

The spin–orbit ground state is doubly degenerate according to our calculation, and thus it cannot be a state with \(\Omega = 0\). This is not in agreement with the assignment of Linton et al.\(^{29}\) who identified the ground state as an \(\Omega = 0\) state. Our calculated first excited state lies 202 cm\(^{-1}\) above the ground state and it is assigned to the \(\Omega = 1\) level at 150 cm\(^{-1}\) reported by Linton.\(^{29}\) We then compute a state at 217 cm\(^{-1}\), which is not seen experimentally. Further states are calculated at 550, 558, and 1013 cm\(^{-1}\), and these are assigned to the 565, 580, and 880 cm\(^{-1}\) experimental excitation energies, respectively.

The chemiionization spectrum recorded for an effusive beam of samarium reacted with atomic oxygen shows a band with a maximum at \((0.02 \pm 0.07)\) eV electron kinetic energy that decreases in intensity to meet the baseline at \((0.28 \pm 0.07)\) eV, the high kinetic energy offset (HKEO).\(^{5,7}\) SmO\(^+\) ions were observed under the same conditions with a mass spectrometer. In the original paper,\(^{5}\) the HKEO value, the enthalpy of the reaction, was calculated from available thermodynamic values as \((0.4 \pm 0.02)\) eV. The shape of the SmO\(^+\) curve in the bond length range of \(4-5\) a.u. is presumably due to a change in nature of the electron distribution and an avoided crossing with another state.

Inspection of an expanded section of Fig. 1, shown in Fig. 2, notably the region where the horizontal line from the Sm + O reactants crosses the SmO\(^+\) potential curve, indicates that this horizontal line first encounters an SmO nonet state at an Sm–O distance of 3.08 a.u. At this bond length, the vibrational wave function at this left-hand turning point overlaps with a SmO\(^+\) potential curve (the SmO\(^+\) \(\Pi\) ground state). Autoionization will therefore occur from this SmO left-hand turning point down to vibrational levels of the lowest SmO\(^+\) state, a \(\Pi\) state. As the SmO nonet autoionizing state and the SmO\(^+\) \(\Pi\) state have essentially the same energy at 3.08 a.u. bond length, the Franck–Condon factors between the neutral state vibrational wave function at the left-hand turning point and the lower vibrational levels of the lowest SmO\(^+\) state.

The energy levels of SmO have been analyzed at a bond distance of 3.08 a.u., where the chemiionization occurs. At this bond distance, the electronic state of SmO that intersects the asymptotic curve of Sm and O atoms is a \(\Pi\) state. This state is lowered in energy by 0.027 eV (226 cm\(^{-1}\)) by the inclusion of spin–orbit coupling and the total splitting of its spin–orbit manifold is 600 cm\(^{-1}\). Inspection of Fig. 2 shows that the Sm + O reaction coordinate (the horizontal dotted line from the Sm + O reagents) crosses a group of SmO nonet states in the bond length region 3.05–3.10 a.u. The first neutral excited state encountered is a \(\Pi\) state as described above. At shorter bond lengths, the Sm + O horizontal dotted line crosses SmO quintet states in the bond length region 2.8–3.0 a.u. and excited septet states in the bond length region 2.6–2.8 a.u. It is interesting that the SmO\(^+\) \(\Pi\) state is autoionizing to take place, as the \(\Delta\Omega = 0\) selection rule is now satisfied.

If the mechanism for the autoionization is considered in a little more detail, for the initial SmO\(^+\) \(\Pi\) state, \(S\) is equal to 4, and possible values of \(S\) in the final state SmO\(^+\) \(\Pi\) or \(\Pi\) are 3 or 2. Therefore the \(\Delta S = 0\) selection rule cannot be satisfied. However, if the \(\Omega\) values in the initial and final state are considered, the \(\Omega\) values for the \(\Pi\) state are 5, 4, 3, 2, 1, 0 and the \(\Omega\) values for the \(\Pi\) state are 13/2, 11/2, 9/2, 7/2, 5/2, 3/2 plus \(e^-\) (\(\Omega = 1/2\)) final state are \(7, 6, 5, 4, 3, 2, 1\). Hence it can be seen that the \(\Delta\Omega = 0\) selection rule can be satisfied by the \(\Omega\) values 5, 4, 3, 2, 1 in the \(\Pi\) state and the final state. It should also be noted that for both the SmO\(^+\) \(\Pi\) and SmO\(^+\) \(\Pi\) states, each \(\Omega\) component is
mixed with other components of the same $\Omega$ value arising from other neutral and ionic states.

IV. CONCLUSIONS

In conclusion, this study is the first theoretical investigation of a metal-plus-oxidant chemiionization reaction. A variety of these reactions involving lanthanides and actinides have been studied experimentally in the gas phase, but this is the first time that one of these reactions has been investigated with quantum chemical methods. The mechanism of the Sm$^+$O chemiionization reaction has been established from the computed potential curves of SmO and SmO$^+$ and the states involved have been identified. The position and shape of the experimental chemi-electron band has also been explained on the basis of these potential curves, and the important role of spin–orbit coupling in the nonet excited state of SmO, in allowing the chemiionization to take place, has been revealed. It is proposed to extend these studies to other metal-plus-oxidant chemiionization reactions.

ACKNOWLEDGMENTS

The authors wish to thank Professor Bjo¨rn Olof Roos for financial support.